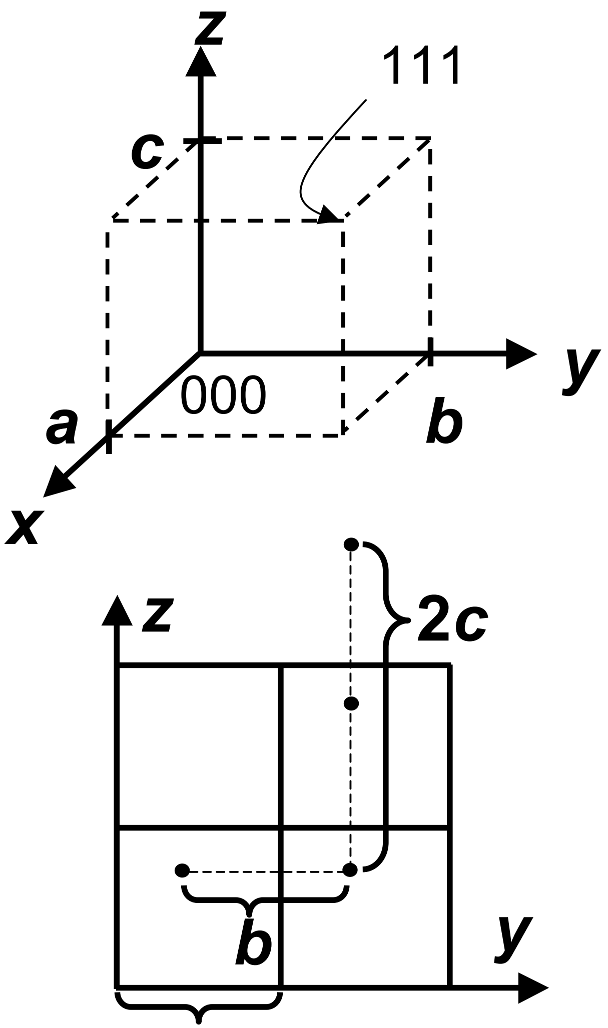
Point coordinates



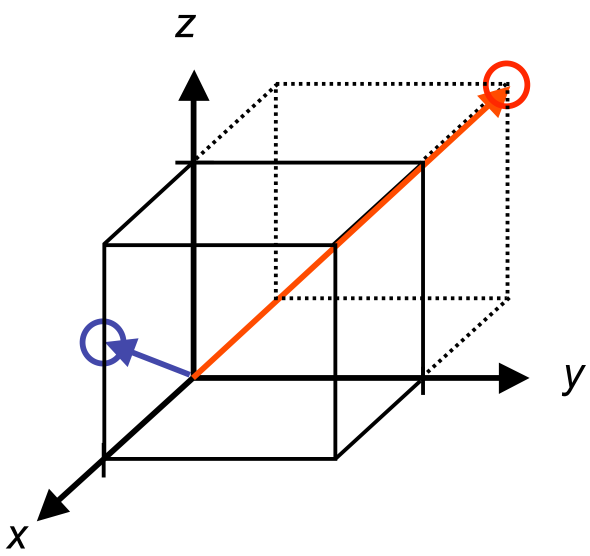
- Point coordinates for unit cell center are

  $a/2$, $b/2$, $c/2$ ½ ½ ½

- Point coordinates for unit cell corner are 111

- Translation: integer multiple of lattice constants → identical position in another unit cell

Directions



Algorithm

1. Vector repositioned (if necessary) to pass through origin.
2. Read off projections in terms of unit cell dimensions *a*, *b*, and *c*
3. Adjust to smallest integer values
4. Enclose in square brackets, no commas

[*xyz*]

ex: 1, 0, ½ => 2, 0, 1 => [201]

-1, 1, 1 => [$\bar{1}$11] where overbar represents a negative index

families of directions <*xyz*>

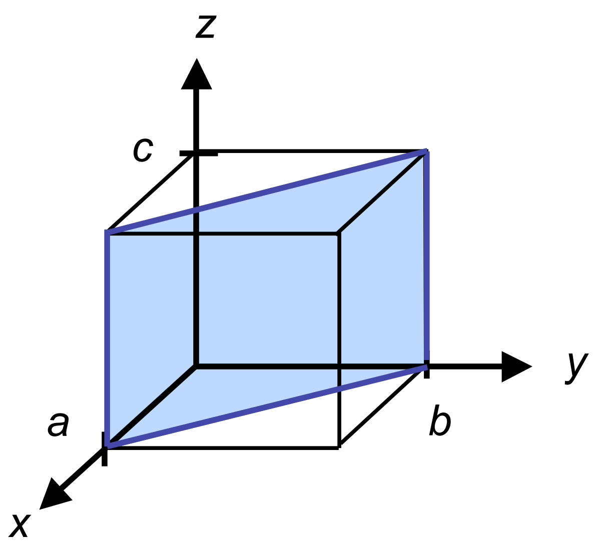
Crystallographic planes

Miller Indices: Reciprocals of the (three) axial intercepts for a plane, cleared of fractions & common multiples. All parallel planes have same Miller indices.

Algorithm

1. Read off intercepts of plane with axes in terms of *a*, *b*, *c*
2. Take reciprocals of intercepts
3. Reduce to smallest integer values
4. Enclose in parentheses, no commas i.e., (*hkl*)

example		*a*	*b*	*c*
1.	Intercepts	1	1	∞
2.	Reciprocals	1/1	1/1	1/∞
3.	Reduction	1	1	0
4.	Miller Indices	(110)		

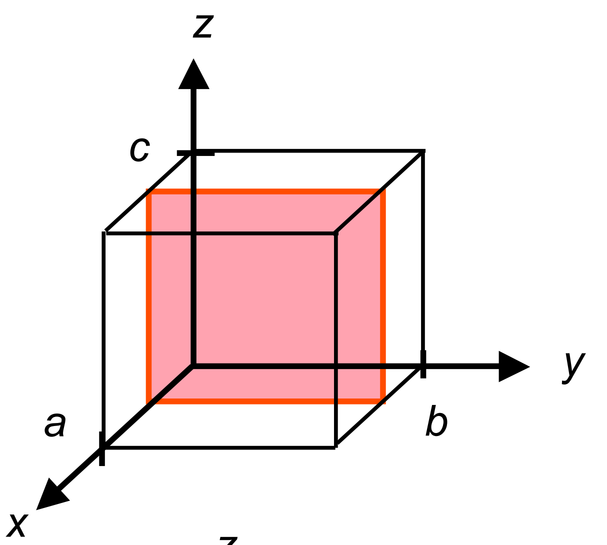


Crystallographic planes

<u>example</u>

		a	b	c
1.	Intercepts	1/2	∞	∞
2.	Reciprocals	1/½	1/∞	1/∞
		2	0	0
3.	Reduction	2	0	0

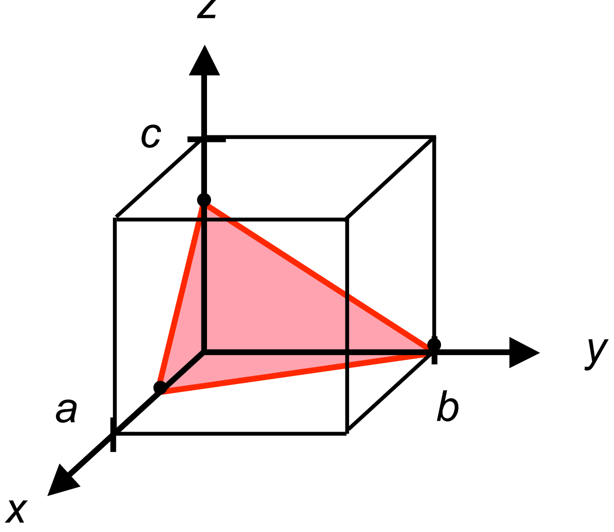
4. Miller Indices (200)



<u>example</u>

		a	b	c
1.	Intercepts	1/2	1	3/4
2.	Reciprocals	1/½	1/1	1/¾
		2	1	4/3
3.	Reduction	6	3	4

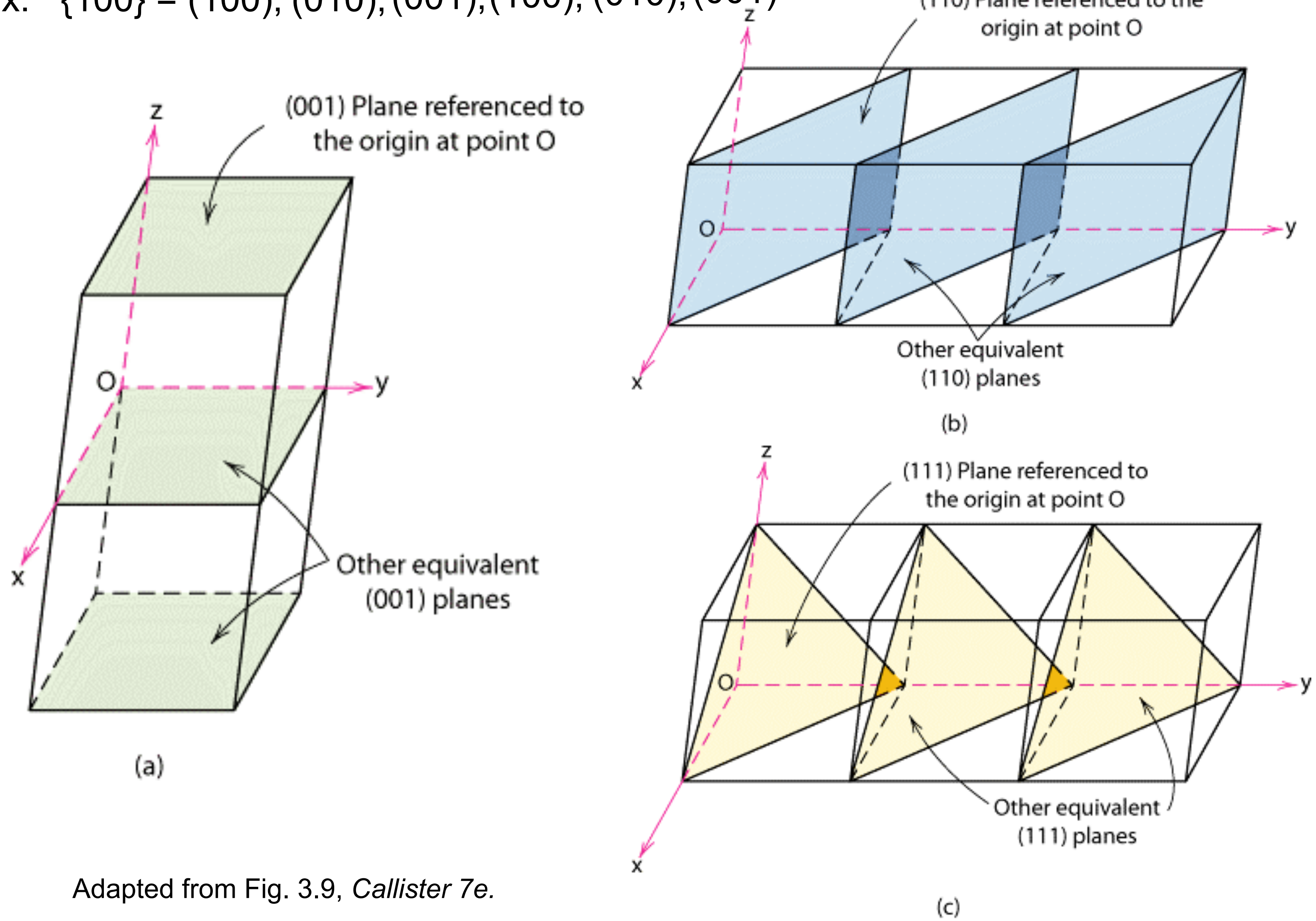
4. Miller Indices (634)



Crystallographic planes

Family of Planes *{hkl}*

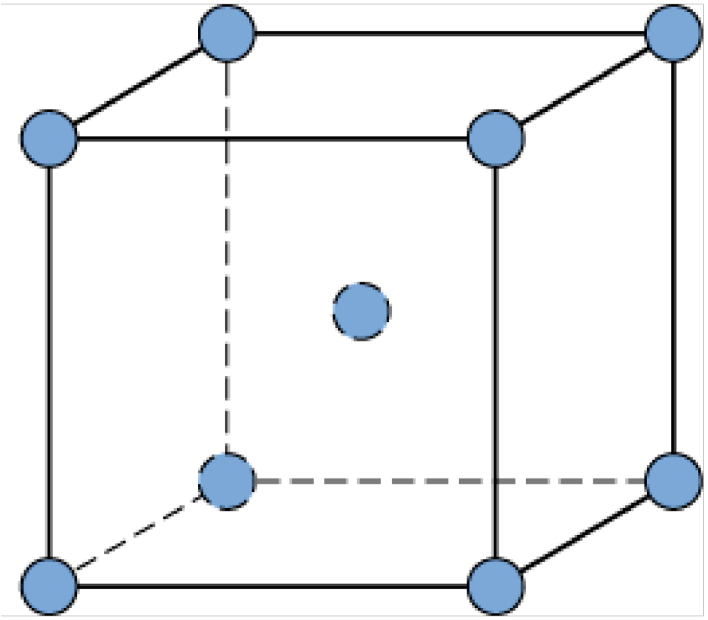
Ex: {100} = (100), (010), (001), $(\bar{1}00)$, $(0\bar{1}0)$, $(00\bar{1})$



Adapted from Fig. 3.9, *Callister 7e.*

Linear density: BCC

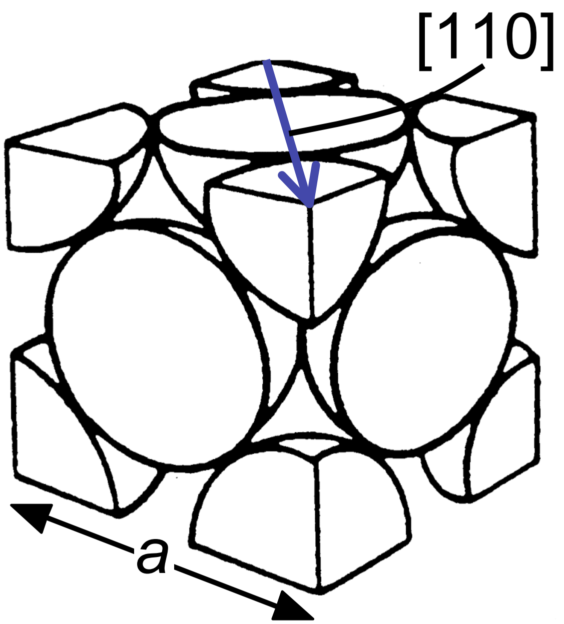
$$\text{Linear Density of Atoms} \equiv \text{LD} = \frac{\text{Number of atoms}}{\text{Unit length of direction vector}}$$



$$[100]: \text{LD} = \frac{1}{a} = \frac{\sqrt{3}}{4R}$$

$$[111]: \text{LD} = \frac{2}{4R} = \frac{1}{2R}$$

Linear density (FCC) and planar densitry



ex: linear density of Al in [110] direction
a = 0.405 nm

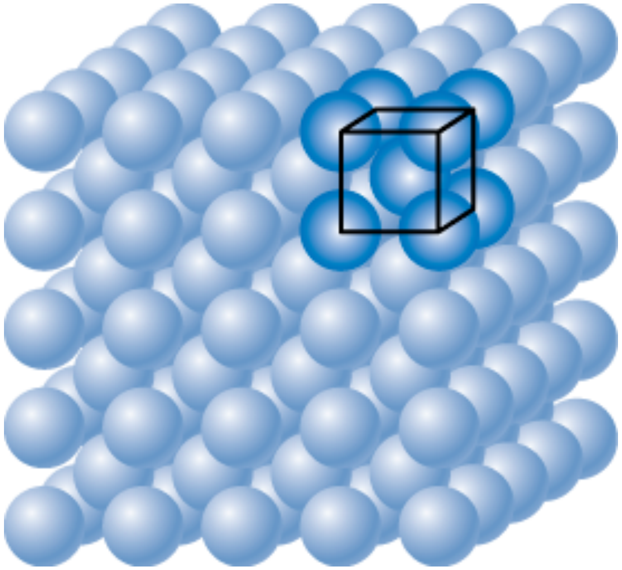
$$LD = \frac{\overset{\#\text{ atoms}}{2}}{\underset{\text{length}}{\sqrt{2}\,a}} = 3.5\text{ nm}^{-1}$$

$$\text{Planar Density of Atoms} \equiv PD = \frac{\text{Number of atoms}}{\text{Area of plane}}$$

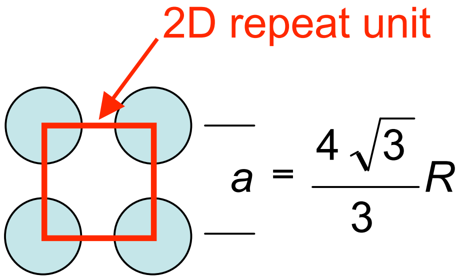
Planar density

- We want to examine the atomic packing of crystallographic planes
- Iron foil can be used as a catalyst. The atomic packing of the exposed planes is important.
  - a) Draw (100) and (111) crystallographic planes for Fe.
  - b) Calculate the planar density for each of these planes.

  R = 0.1241 nm and Fe has a BCC structure at room temperature



Adapted from Fig. 3.2(c), *Callister 7e.*

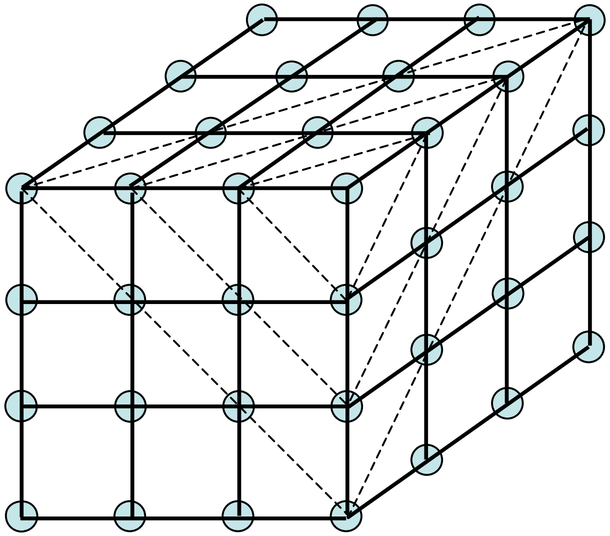


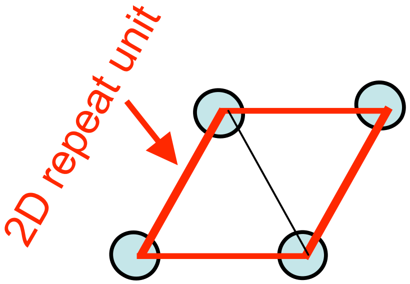
$$a = \frac{4\sqrt{3}}{3}R$$

$$\text{Planar Density} = \frac{\dfrac{\text{atoms}}{\text{2D repeat unit}}}{\dfrac{\text{area}}{\text{2D repeat unit}}} = \frac{1}{a^2} = \frac{1}{\left(\dfrac{4\sqrt{3}}{3}R\right)^2}$$

$$= \boxed{1.2 \times 10^{19} \frac{\text{atoms}}{\text{m}^2}}$$

Planar density





$$\text{area} = \sqrt{2}\,ah = \sqrt{3}\,a^{2} = \sqrt{3}\left(\frac{4\sqrt{3}}{3}R\right)^{2} = \frac{16\sqrt{3}}{3}R^{2}$$

$$\text{Planar Density} = \frac{\dfrac{\text{atoms}}{\text{2D repeat unit}}}{\dfrac{\text{area}}{\text{2D repeat unit}}} = \frac{1}{\dfrac{16\sqrt{3}}{3}R\ 2} = 7.0\,\frac{\text{atoms}}{\text{nm}^{2}} = \boxed{0.70 \times 10^{19}\,\frac{\text{atoms}}{\text{m}^{2}}}$$

Single crystal vs. Polycrystalline structures

Single crystals: Atoms all have the same arrangement throughout.

Polycrystalline: Many crystals put together.

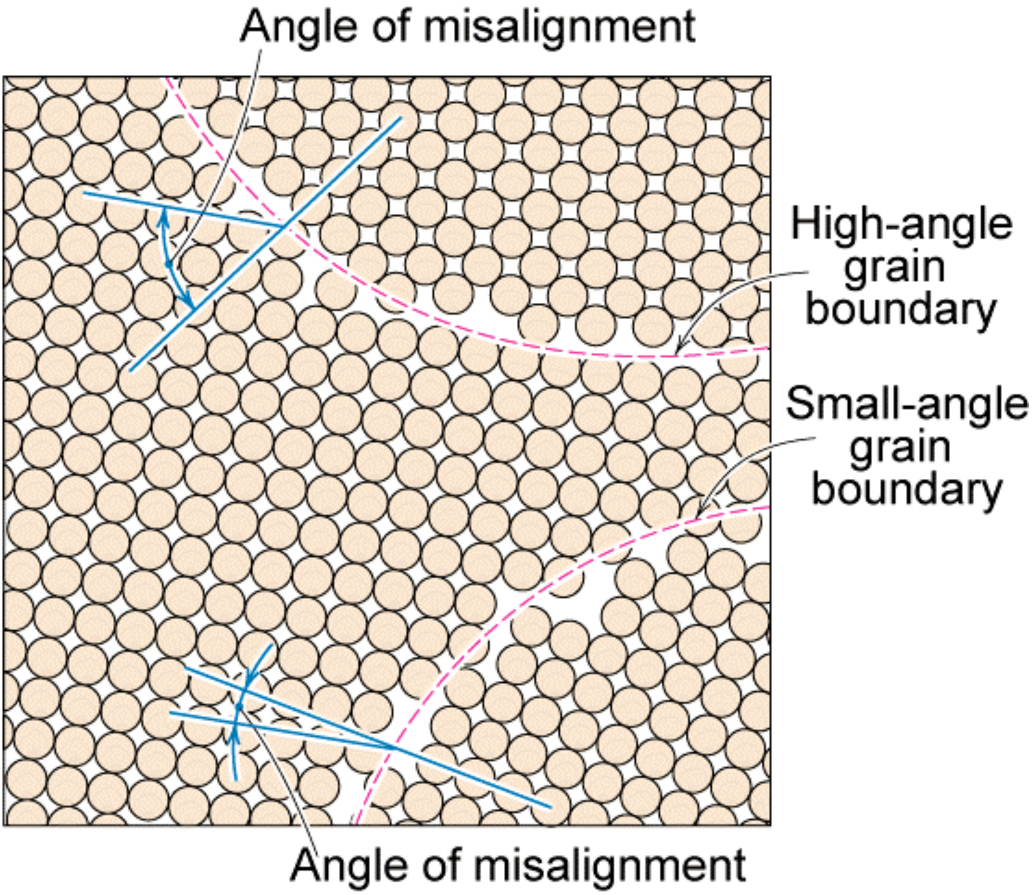
Adapted from Fig. K, color inset pages of *Callister 5e*. (Fig. K is courtesy of Paul E. Danielson, Teledyne Wah Chang Albany)

- Nb-Hf-W plate with an electron beam weld.
- Each "grain" is a single crystal.
- If grains are randomly oriented, overall component properties are not directional.
- Grain sizes typ. range from 1 nm to 2 cm (i.e., from a few to millions of atomic layers).

Polycrystalline structures

Grain Boundaries

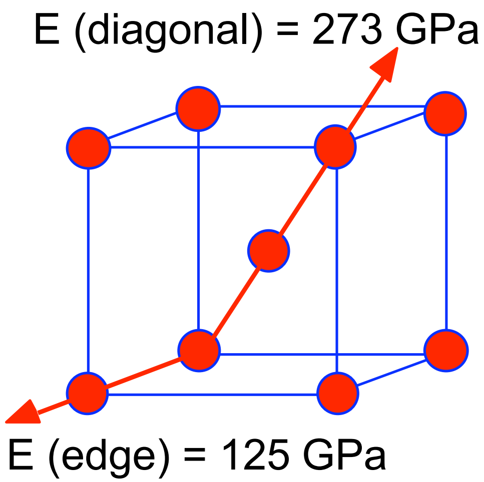
- regions between crystals
- transition from lattice of one region to that of the other
- slightly disordered
- low density in grain boundaries
  - o high mobility
  - o high diffusivity
  - o high chemical reactivity



Adapted from Fig. 4.7, *Callister 7e.*

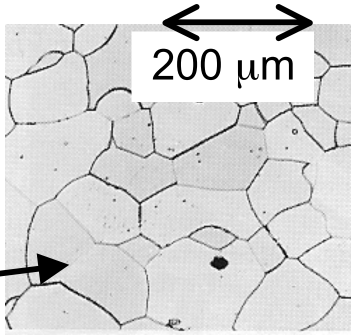
Single crystal vs. Polycrystalline structures

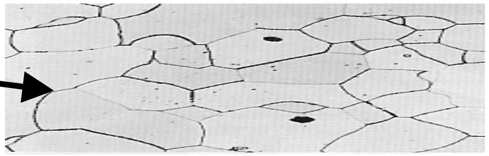
- Single Crystals
  - -Properties vary with direction: anisotropic.
  - -Example: the modulus of elasticity (E) in BCC iron:



Data from Table 3.3, *Callister 7e*. (Source of data is R.W. Hertzberg, *Deformation and Fracture Mechanics of Engineering Materials*, 3rd ed., John Wiley and Sons, 1989.)

- Polycrystals
  - -Properties may/may not vary with direction.
  - -If grains are randomly oriented: isotropic. ($E_{\text{poly iron}}$ = 210 GPa)
  - -If grains are textured, anisotropic.



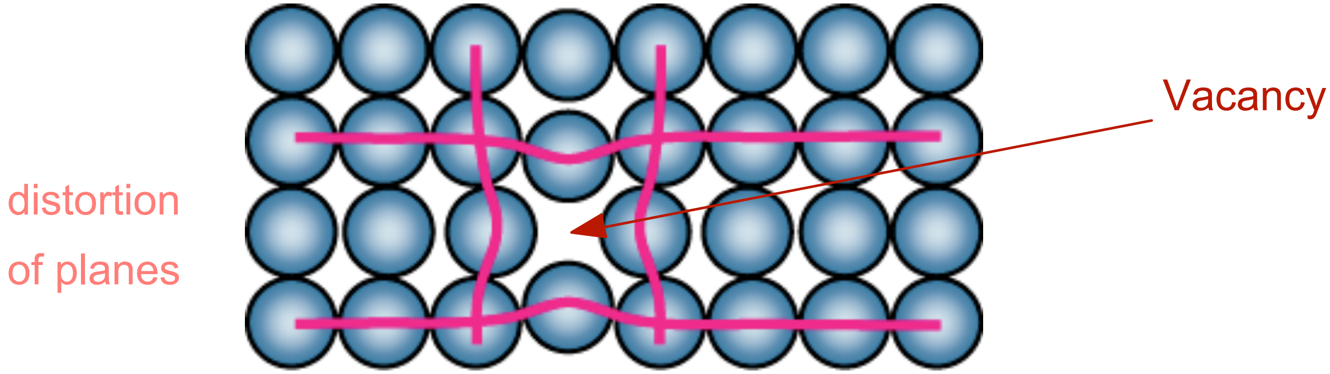


Adapted from Fig. 4.14(b), *Callister 7e*. (Fig. 4.14(b) is courtesy of L.C. Smith and C. Brady, the National Bureau of Standards, Washington, DC [now the National Institute of Standards and Technology, Gaithersburg, MD].)

Point defects

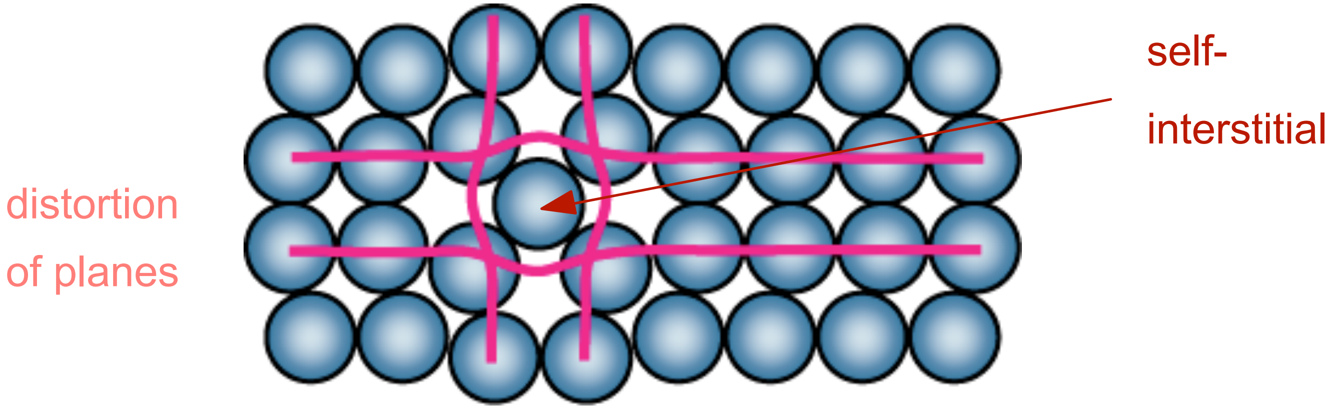
- Vacancies:

  -vacant atomic sites in a structure.



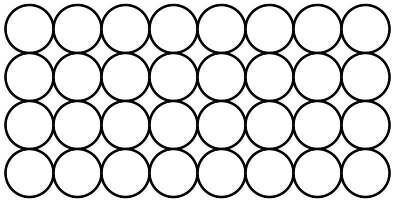
- Self-Interstitials:

  -"extra" atoms positioned between atomic sites.



Equilibrium concentration of defects

- Equilibrium concentration varies with temperature



Each lattice site is a potential vacancy

No. of defects

Activation energy

$$\frac{N_v}{N} = \exp\left(\frac{-Q_v}{k\ T}\right)$$

No. of potential defect sites.

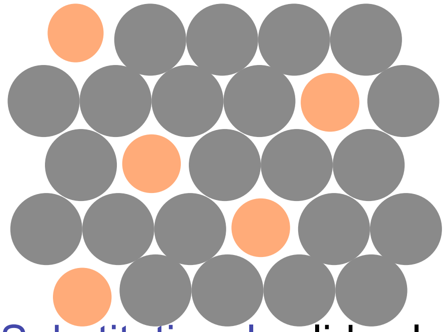
Temperature

Boltzmann's constant

Point defects in alloys

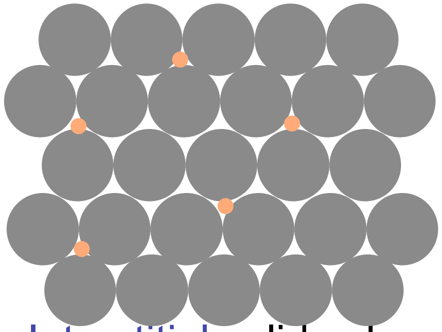
Two outcomes if impurity (B) added to host (A):

- Solid solution of B in A (i.e., random dist. of point defects)



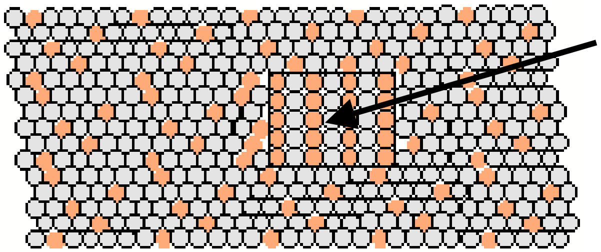
Substitutional solid soln.
(e.g., Cu in Ni)

OR



Interstitial solid soln.
(e.g., C in Fe)

- Solid solution of B in A plus particles of a new phase (usually for a larger amount of B)



Second phase particle
--different composition
--often different structure.

Imperfections of solids

Conditions for substitutional solid solution (S.S.)

W. Hume – Rothery rule

1. Δr (atomic radius) < 15%
2. Proximity in periodic table
   i.e., similar electronegativities
3. Same crystal structure for pure metals
4. Valency
   All else being equal, a metal will have a greater tendency to dissolve a metal of higher valency than one of lower valency

1. Would you predict more Al or Ag to dissolve in Zn?

2. More Zn or Al in Cu?

Element	*Atomic Radius (nm)*	*Crystal Structure*	*Electro-nega-tivity*	*Valence*
Cu	0.1278	FCC	1.9	+2
C	0.071			
H	0.046			
O	0.060			
Ag	0.1445	FCC	1.9	+1
Al	0.1431	FCC	1.5	+3
Ni	0.1246	FCC	1.8	+2
Zn	0.1332	HCP	1.6	+2

Table on p. 106, *Callister 7e.*

Line defects

- Linear Defects (Dislocations)
    - Are one-dimensional defects around which atoms are misaligned
- Edge dislocation:
    - extra half-plane of atoms inserted in a crystal structure
    - **b** ⊥ to dislocation line
- Screw dislocation:
    - spiral planar ramp resulting from shear deformation
    - **b** || to dislocation line

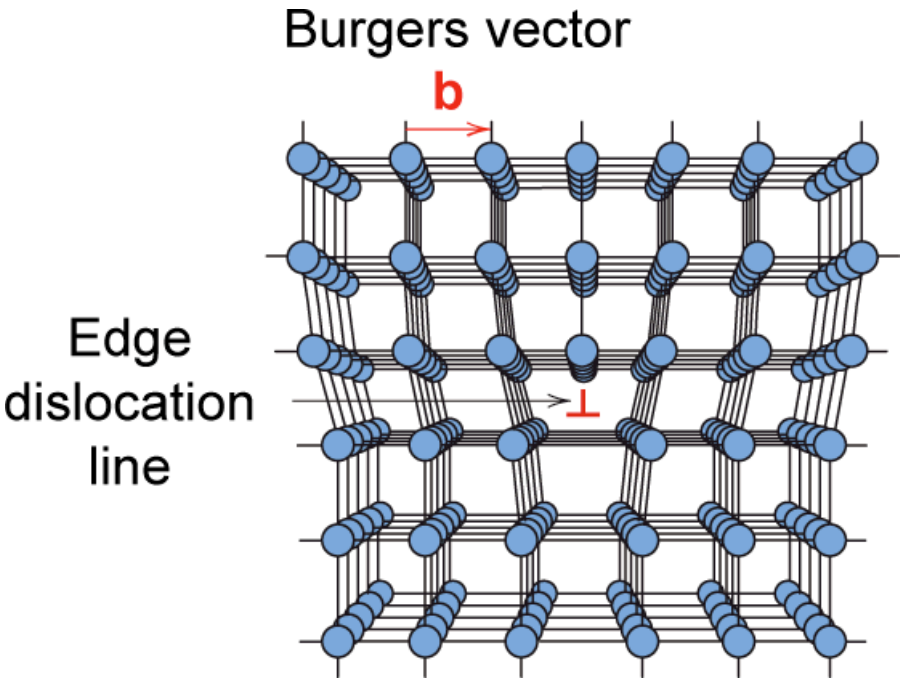
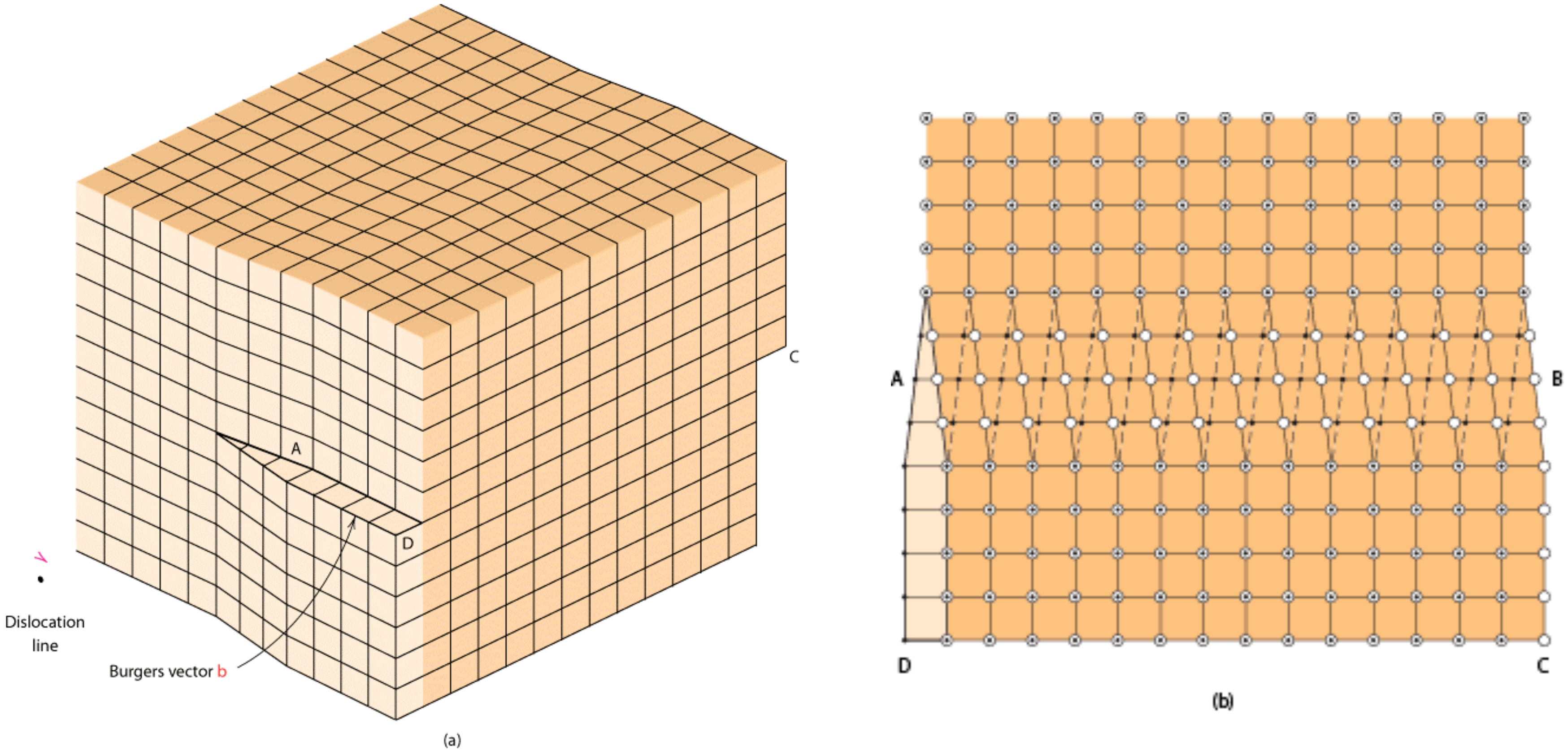


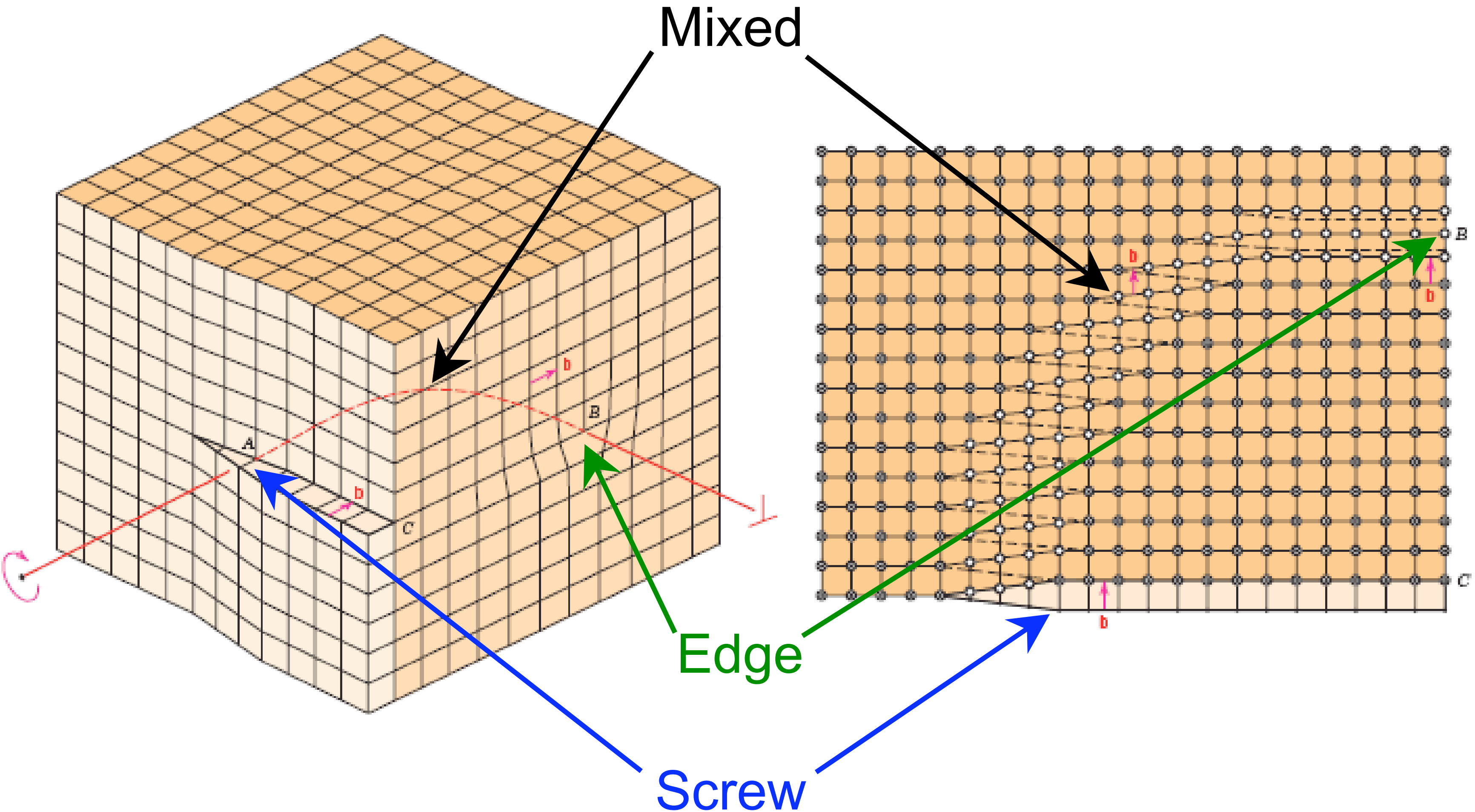
Fig. 4.3, *Callister 7e.*

Screw defects



Adapted from Fig. 4.4, *Callister 7e.*

Mixed defects



Mixed

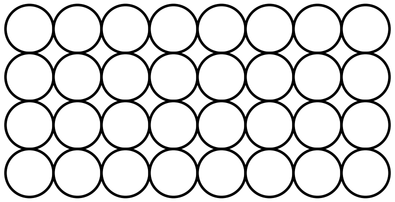
Edge

Screw

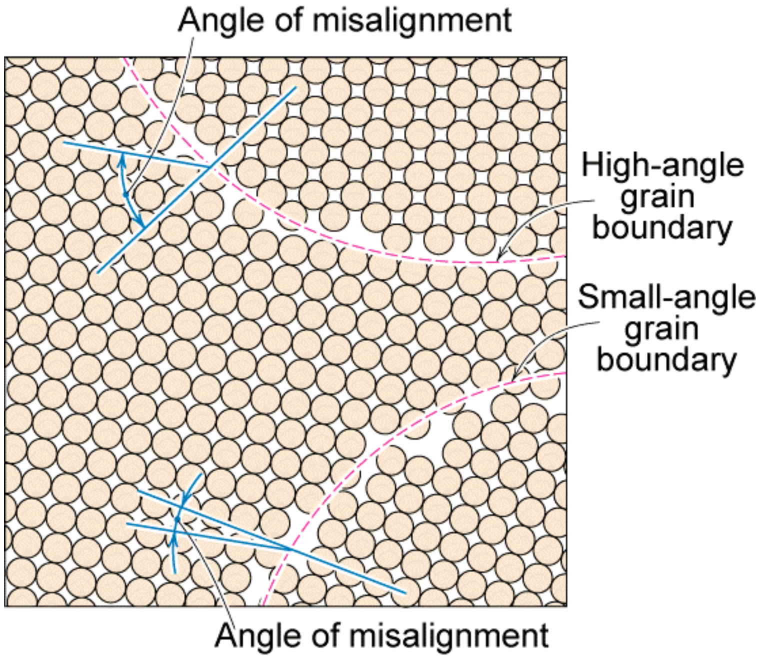
Adapted from Fig. 4.5, *Callister 7e.*

Planar defects

- External defects
  - On the surface
  - Unsatisfied bonds



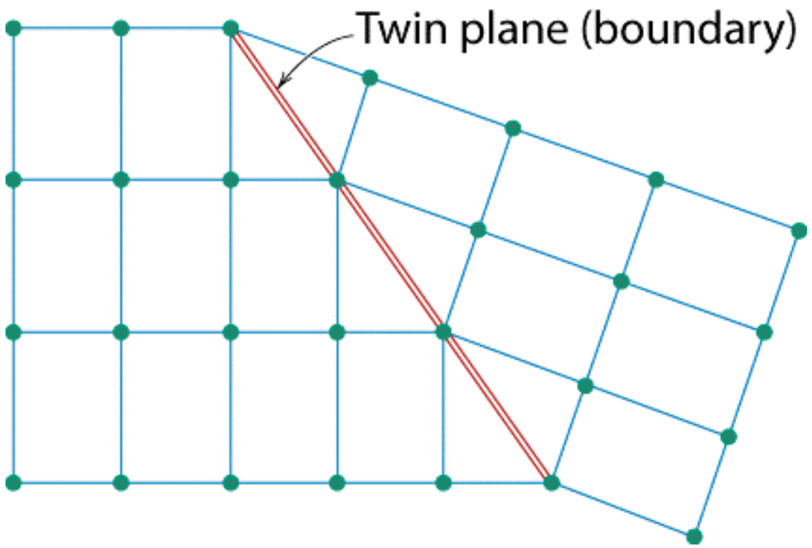
- Internal defects
  - Grain boundaries



Planar defects

- Twin boundary (plane)

  Essentially a reflection of atom positions across the twin plane.



Adapted from Fig. 4.9, *Callister 7e.*

- Stacking faults

  For FCC metals an error in ABCABC packing sequence
  Ex: ABCABABC