Electrical conductivity

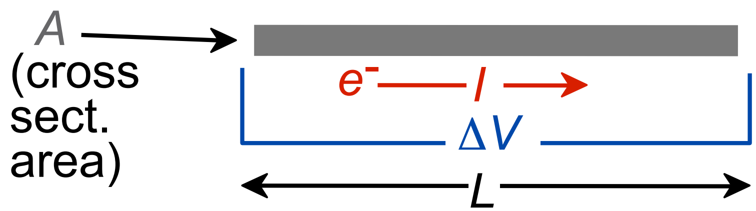
- Ohm's Law:

$$\Delta V = I\, R$$

voltage drop (volts = J/C)
C = Coulomb

current (amps = C/s)

resistance (Ohms)



- Resistivity, ρ and Conductivity, σ:

  -- geometry-independent forms of Ohm's Law

  -- Resistivity is a material property & is independent of sample

Conductivity: comparison

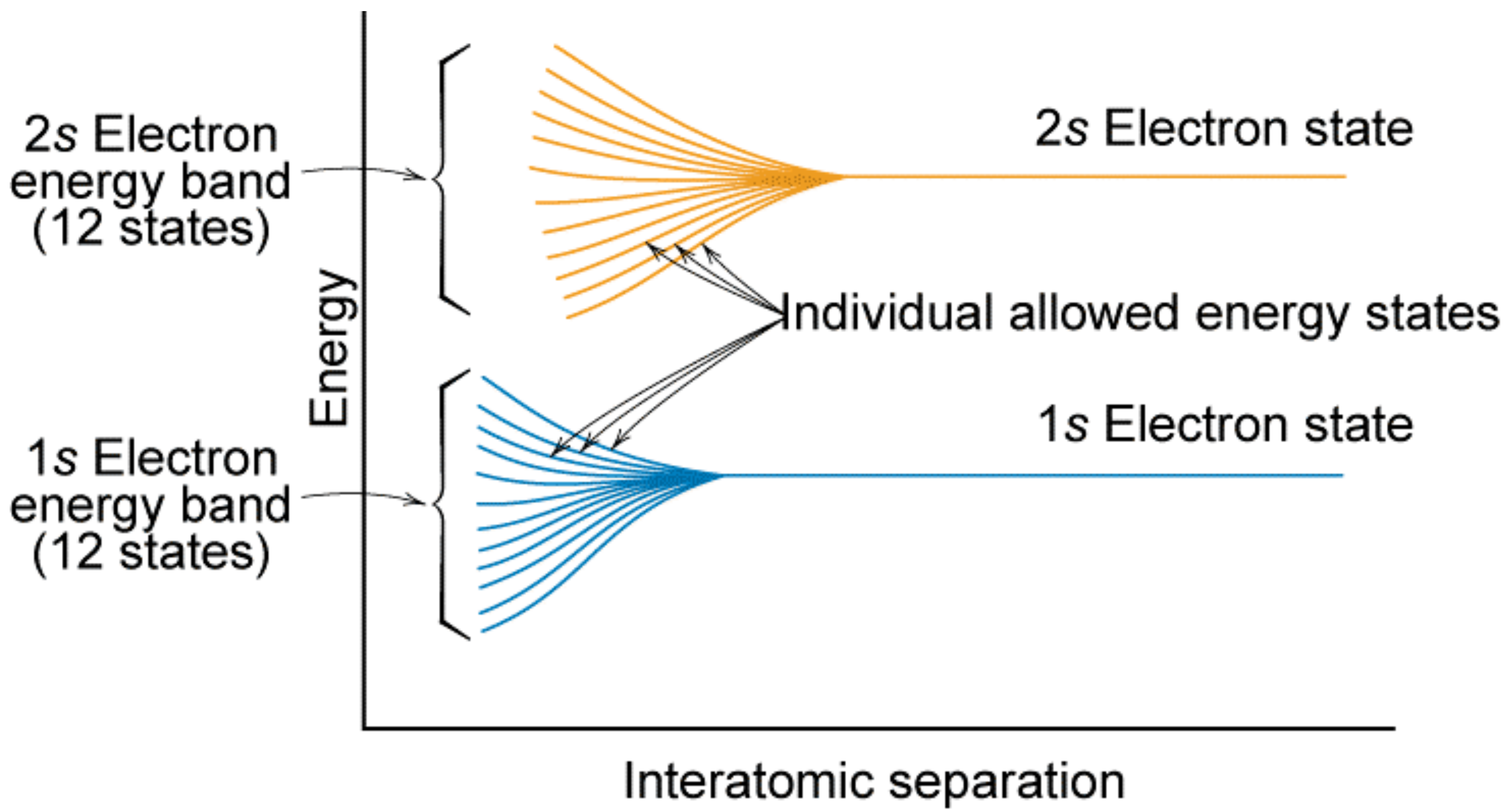
- Room T values $(\text{Ohm-m})^{-1} = (\Omega - \text{m})^{-1}$

METALS	conductors	CERAMICS	
Silver	6.8×10^{7}	Soda-lime glass	10^{-10}-10^{-11}
Copper	6.0×10^{7}	Concrete	10^{-9}
Iron	1.0×10^{7}	Aluminum oxide	$<10^{-13}$

SEMICONDUCTORS		POLYMERS	
Silicon	4×10^{-4}	Polystyrene	$<10^{-14}$
Germanium	2×10^{0}	Polyethylene	10^{-15}-10^{-17}
GaAs	10^{-6}		
	semiconductors		insulators

Selected values from Tables 18.1, 18.3, and 18.4, *Callister 7e*.

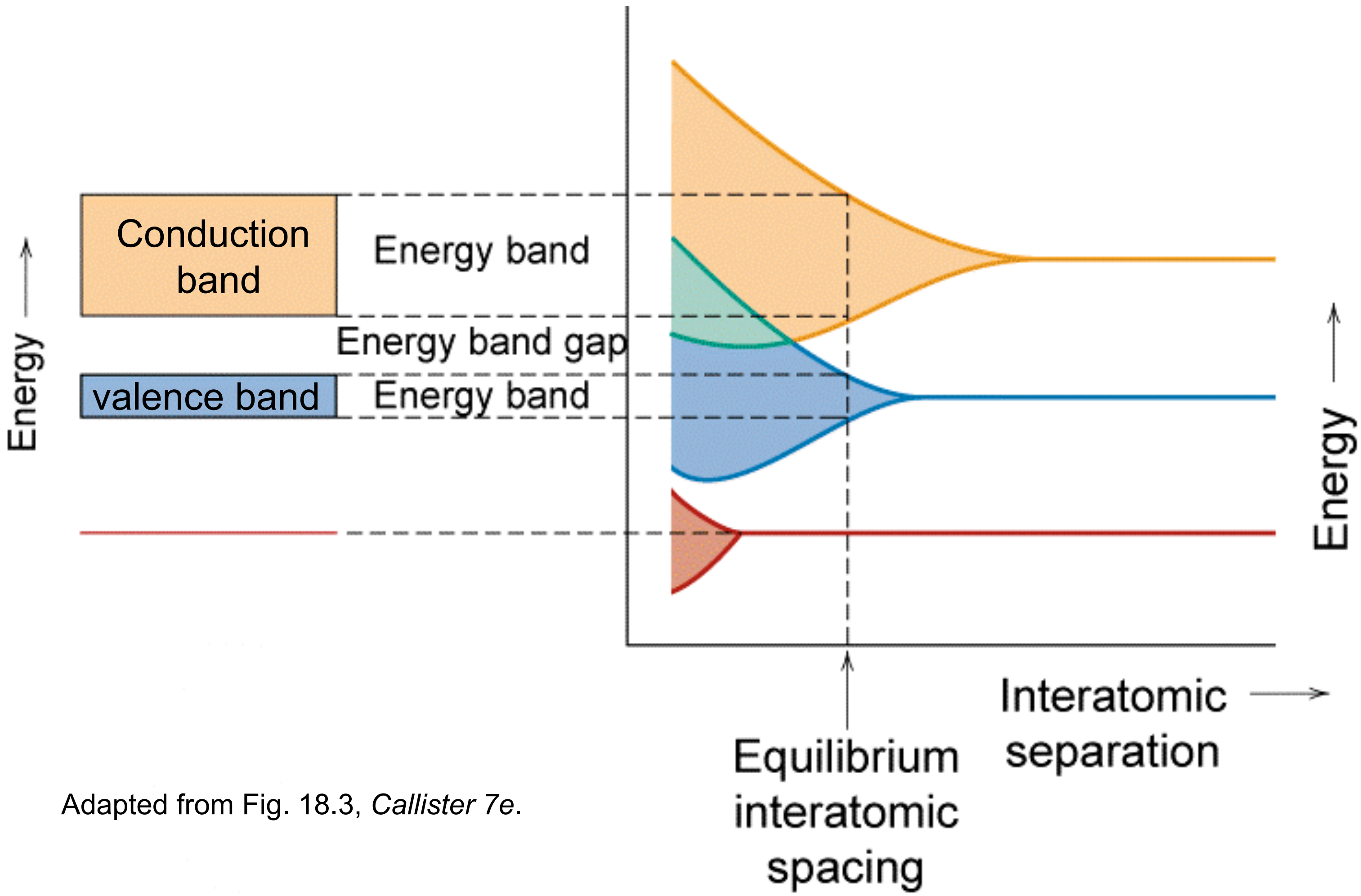
Electronic band structure



Adapted from Fig. 18.2, *Callister 7e*.

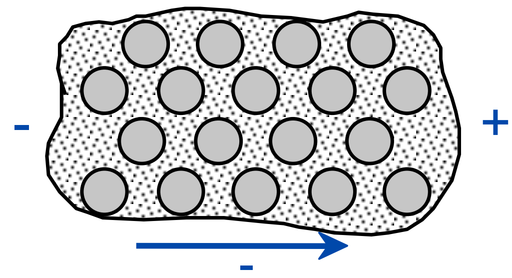
Electronic band structure

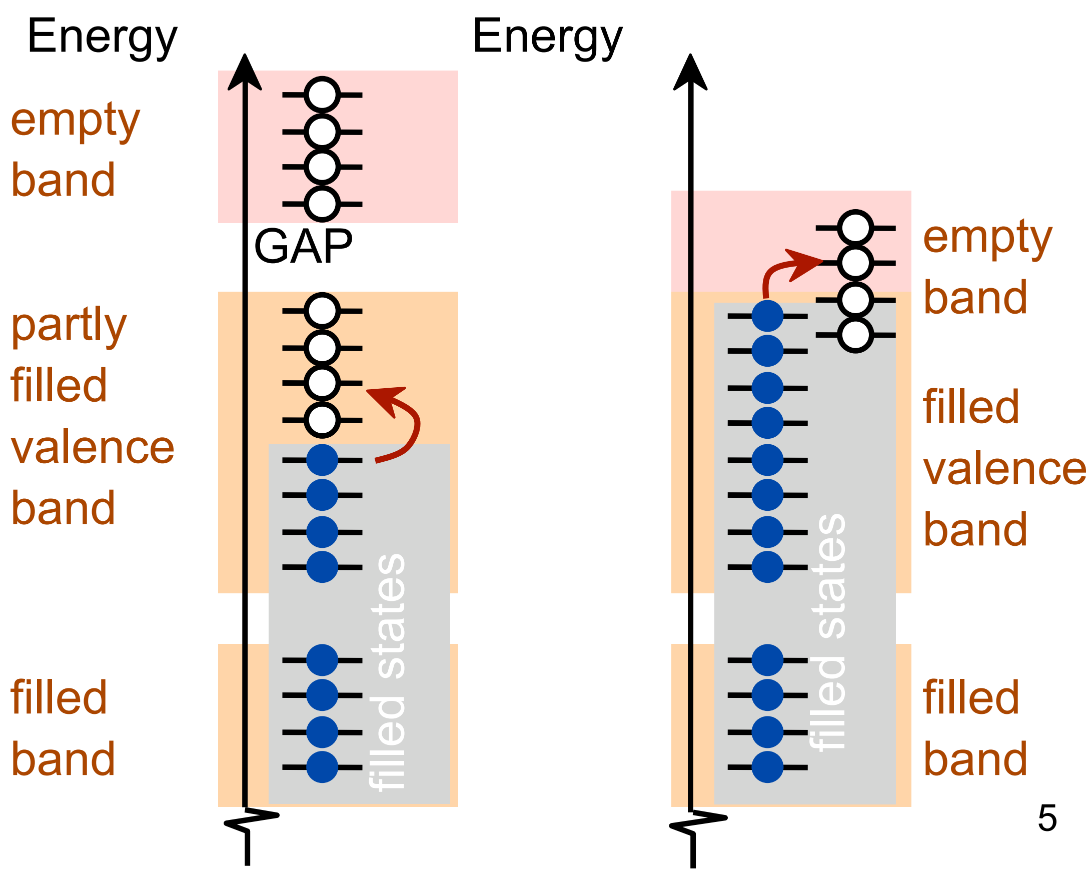
- Valence band – filled – highest occupied energy levels
- Conduction band – empty – lowest unoccupied energy levels



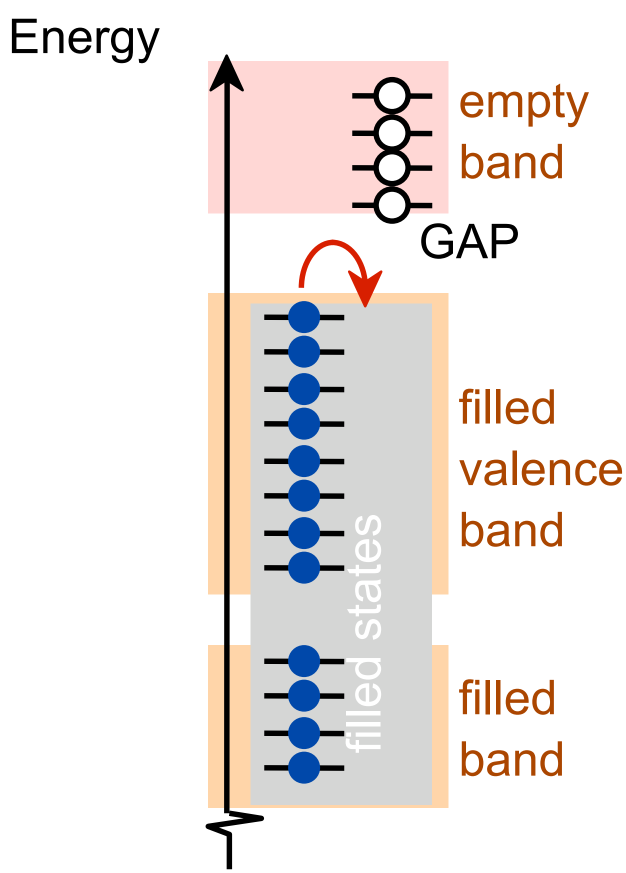
Adapted from Fig. 18.3, *Callister 7e*.

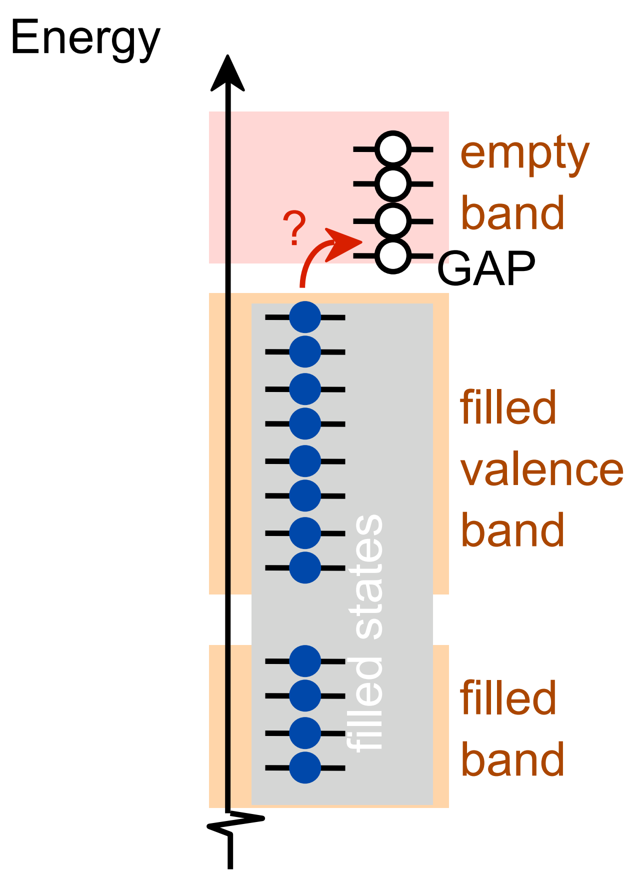
Conduction and electron transport





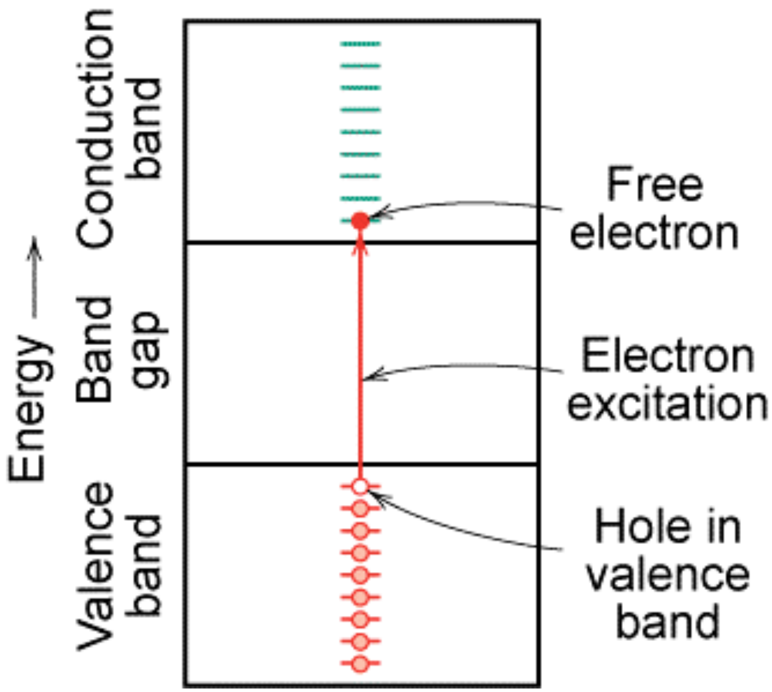
Energy states: Insulators & semiconductors





Charge carriers

Adapted from Fig. 18.6 (b), *Callister 7e*.



Two charge carrying mechanisms

Electron – negative charge

Hole – equal & opposite positive charge

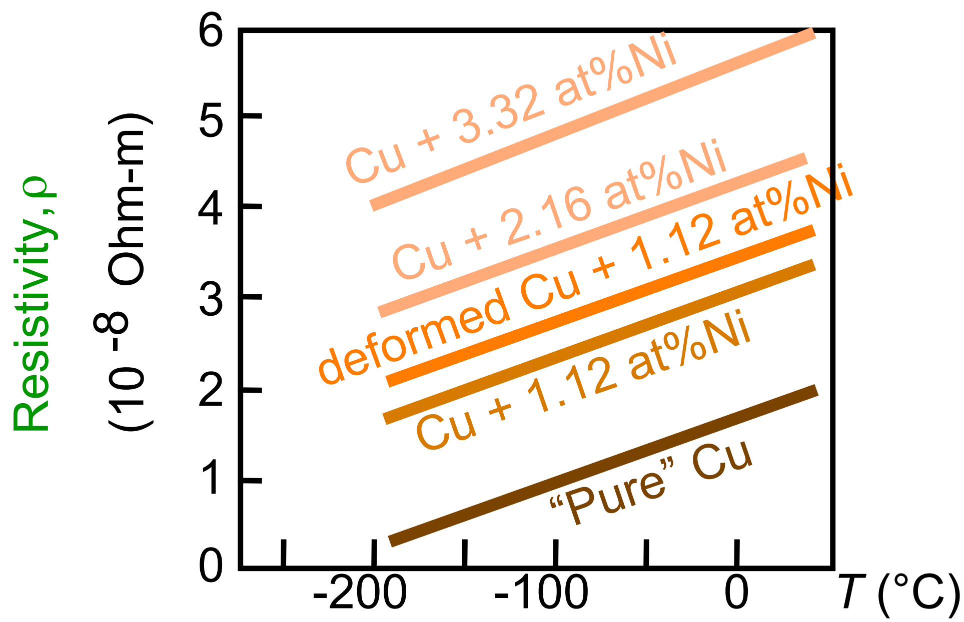
Move at different speeds - drift velocity

Higher temp. promotes more electrons into the conduction band

∴ $\sigma\uparrow$ as $T\uparrow$

Electrons scattered by impurities, grain boundaries, etc.

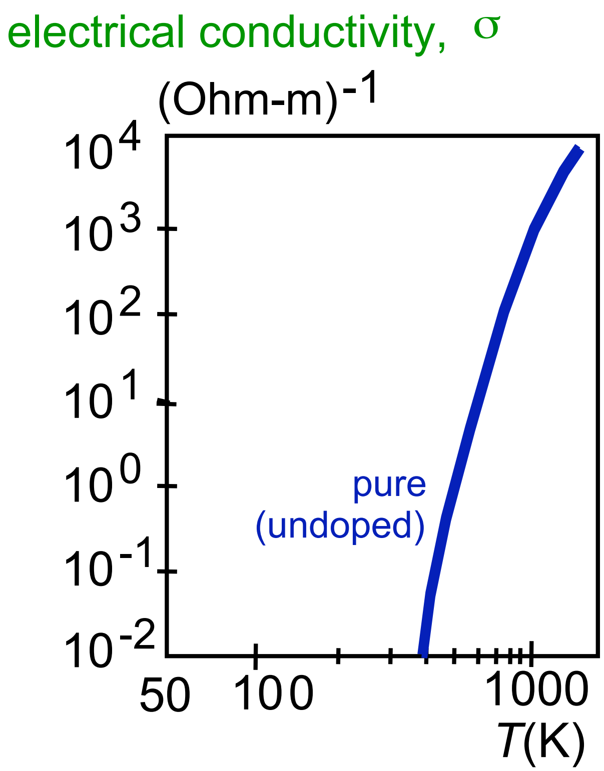
Charge carriers



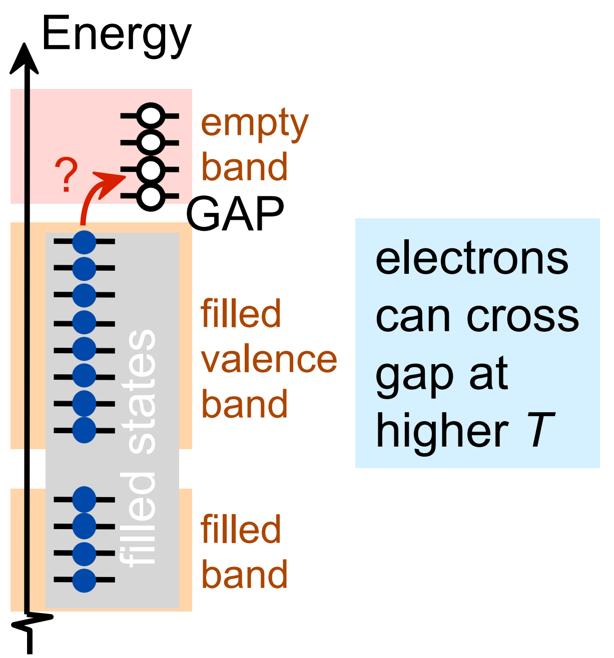
Adapted from Fig. 18.8, *Callister 7e*. (Fig. 18.8 adapted from J.O. Linde, *Ann. Physik* **5**, p. 219 (1932); and C.A. Wert and R.M. Thomson, *Physics of Solids*, 2nd ed., McGraw-Hill Book Company, New York, 1970.)

Pure semiconductors: Conductivity vs T

- Data for Pure Silicon:
  -- σ increases with T
  -- opposite to metals



Adapted from Fig. 19.15, *Callister 5e*. (Fig. 19.15 adapted from G.L. Pearson and J. Bardeen, *Phys. Rev.* **75**, p. 865, 1949.)

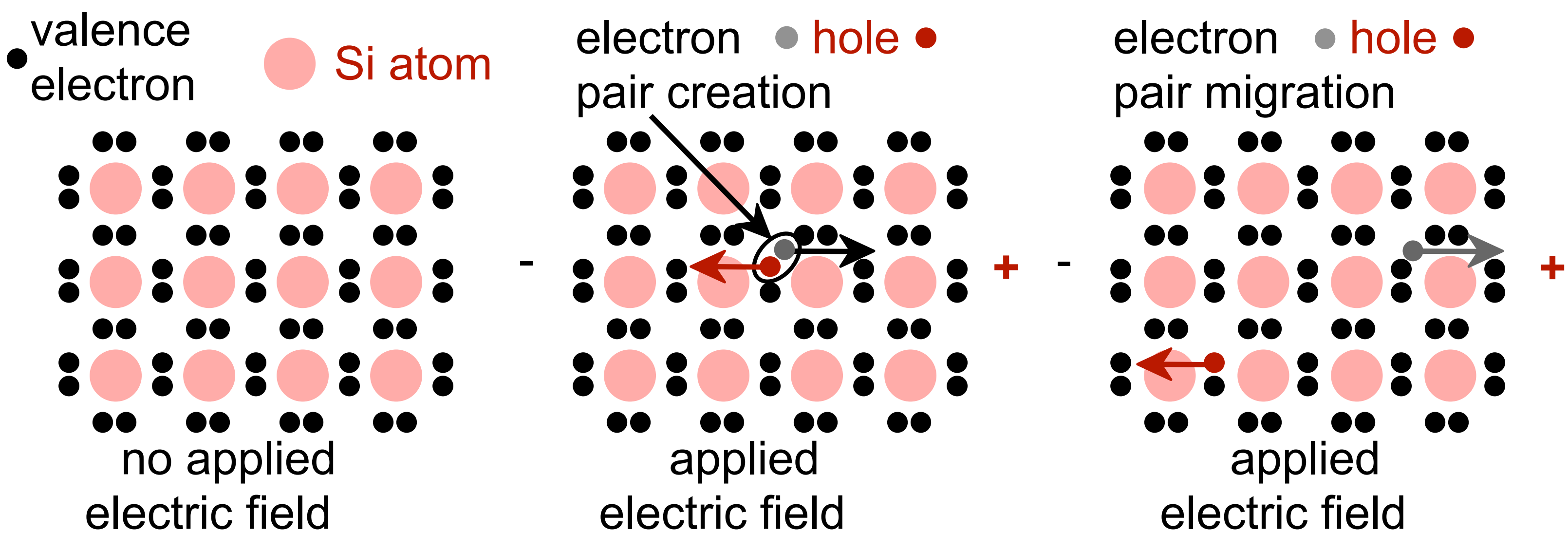


material	band gap (eV)
Si	1.11
Ge	0.67
GaP	2.25
CdS	2.40

Selected values from Table 18.3, *Callister 7e*.

Conduction in terms of electron and hole migration

- Concept of electrons and holes:



Adapted from Fig. 18.11, *Callister 7e*.

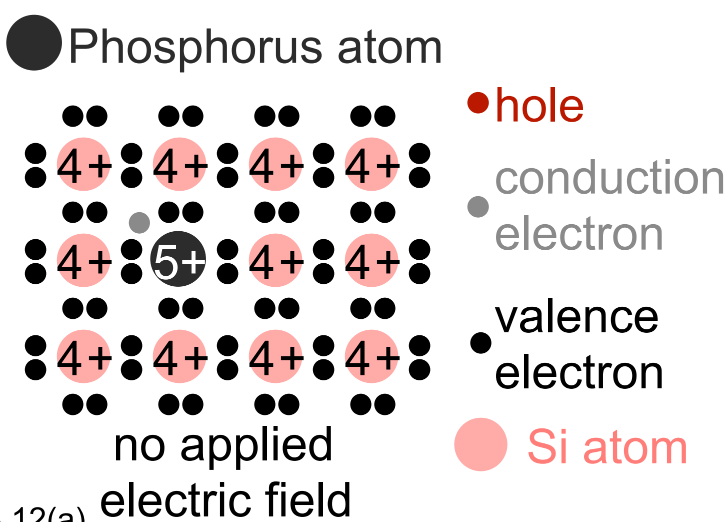
Intrinsic vs extrinsic conduction

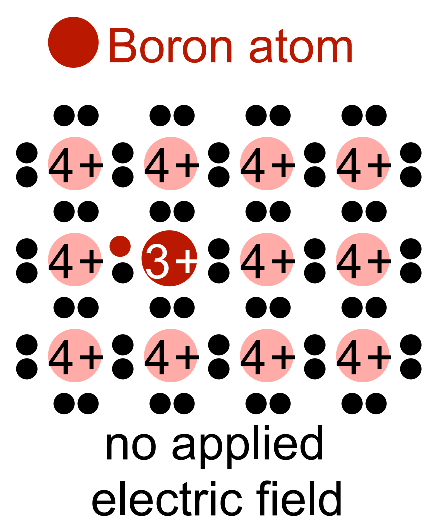
- Intrinsic:
  # electrons = # holes ($n = p$)
  --case for pure Si

- Extrinsic:
  --$n \neq p$
  --occurs when impurities are added with a different
  # valence electrons than the host (e.g., Si atoms)

- *n*-type Extrinsic: ($n >> p$)
- *p*-type Extrinsic: ($p >> n$)

$$\sigma \approx n|e|\mu_e$$



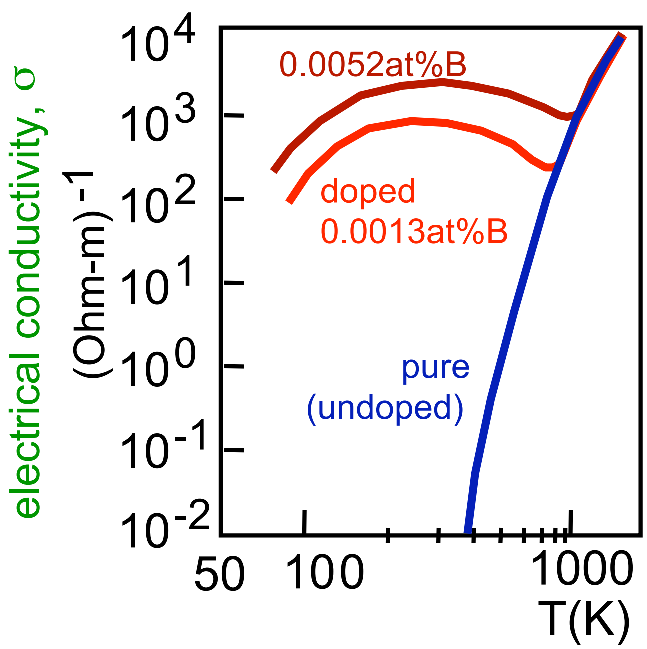


$$\sigma \approx p|e|\mu_h$$

Adapted from Figs. 18.12(a) & 18.14(a), *Callister 7e*.

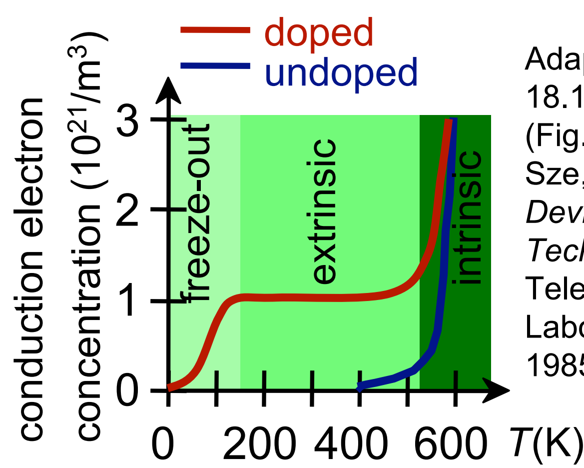
Doped semiconductor: conductivity vs. T

- Data for Doped Silicon:
  -- σ increases with doping
  -- reason: imperfection sites lower the activation energy to produce mobile electrons.



Adapted from Fig. 19.15, *Callister 5e*. (Fig. 19.15 adapted from G.L. Pearson and J. Bardeen, *Phys. Rev.* **75**, p. 865, 1949.)

- Comparison: intrinsic vs extrinsic conduction...
  -- extrinsic doping level: $10^{21}/m^3$ of a *n*-type donor impurity (such as P).
  -- for $T < 100$ K: "freeze-out", thermal energy insufficient to excite electrons.
  -- for 150 K $< T <$ 450 K: "extrinsic"
  -- for $T >> 450$ K: "intrinsic"



Adapted from Fig. 18.17, *Callister 7e*. (Fig. 18.17 from S.M. Sze, *Semiconductor Devices, Physics, and Technology*, Bell Telephone Laboratories, Inc., 1985.)

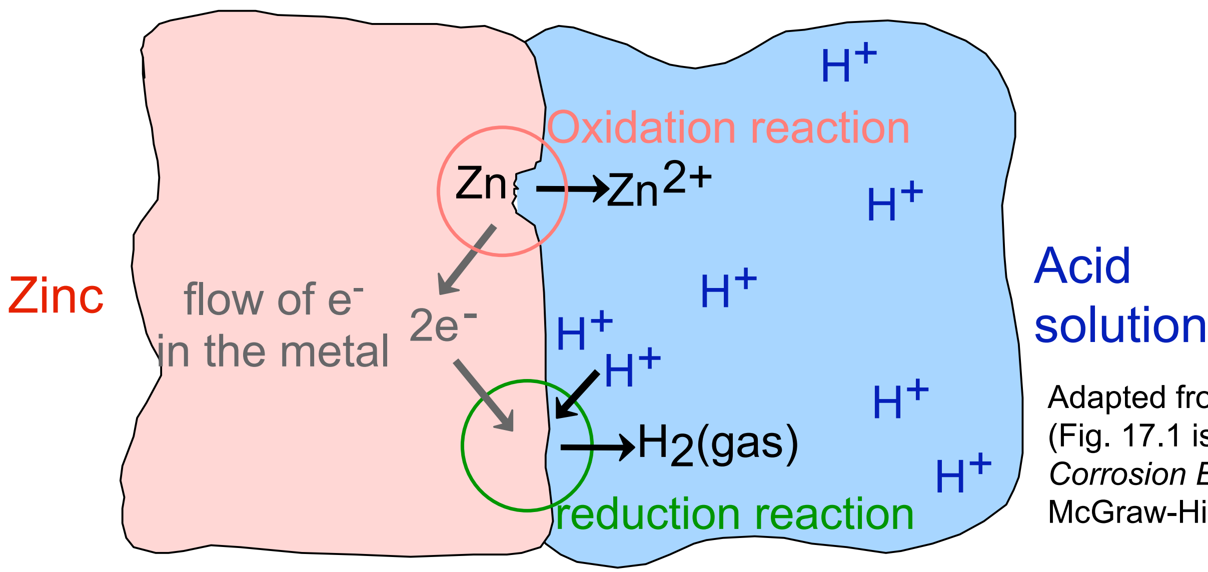
Doped semiconductor: conductivity vs. T

Intrinsic Conductivity

$$\sigma = n|e|\mu_e + p|e|\mu_e$$

Corrosion of zinc in acid

- Two reactions are necessary:
  -- oxidation reaction:
  -- reduction reaction:

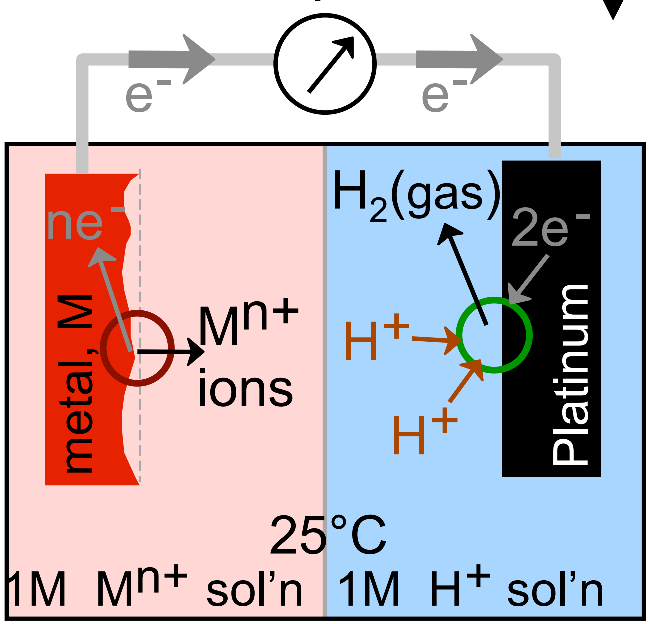


Adapted from Fig. 17.1, *Callister 7e*.
(Fig. 17.1 is from M.G. Fontana, *Corrosion Engineering*, 3rd ed., McGraw-Hill Book Company, 1986.)

Standard hydrogen (EMF) test

- Two outcomes:

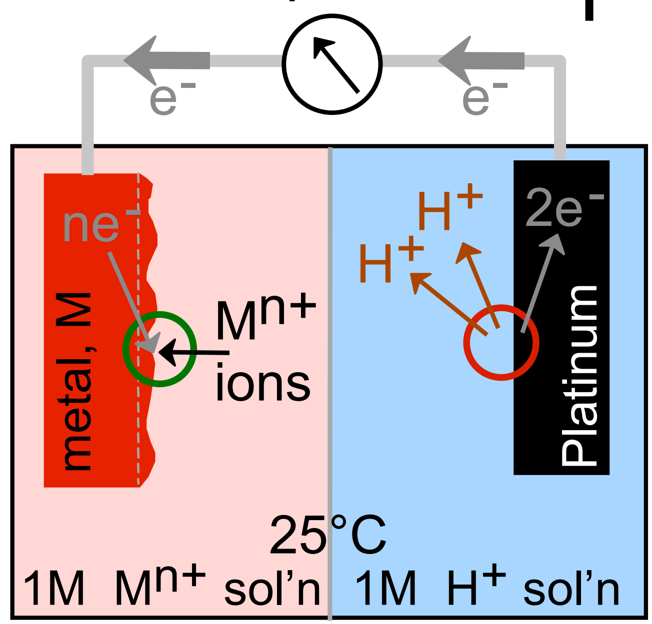
--Metal sample mass ↓



--Metal is the anode (-)

$V^o_{metal} < 0$ (relative to Pt)

--Metal sample mass ↑



--Metal is the cathode (+)

$V^o_{metal} > 0$ (relative to Pt)

Standard Electrode Potential

Adapted from Fig. 17.2, *Callister 7e*.

Standard EMF series

- EMF series

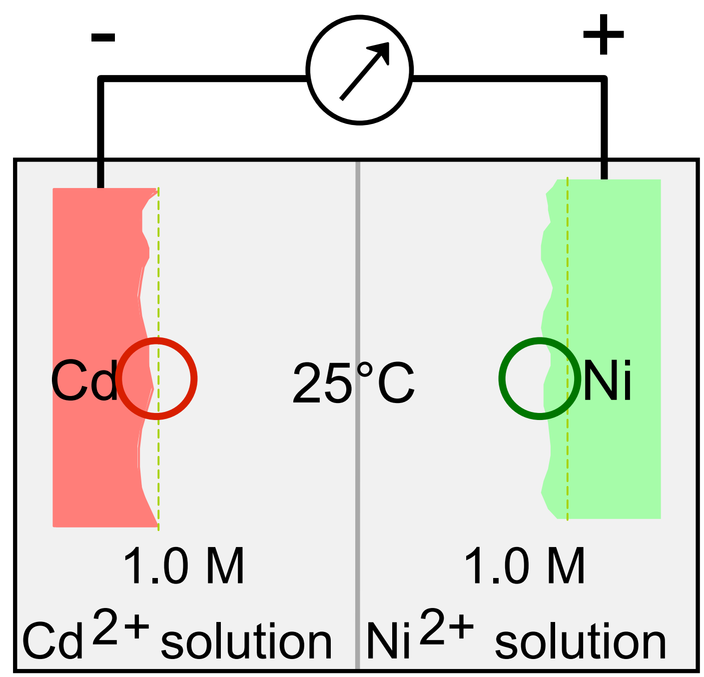
more cathodic ↑ / more anodic ↓

metal	V^o_{metal}
Au	+1.420 V
Cu	+0.340
Pb	- 0.126
Sn	- 0.136
Ni	- 0.250
Co	- 0.277
Cd	- 0.403
Fe	- 0.440
Cr	- 0.744
Zn	- 0.763
Al	- 1.662
Mg	- 2.363
Na	- 2.714
K	- 2.924

$\Delta V^o =$ 0.153V (between Ni and Cd)

Data based on Table 17.1, *Callister 7e.*

- Metal with smaller V^o_{metal} corrodes.
- Ex: Cd-Ni cell

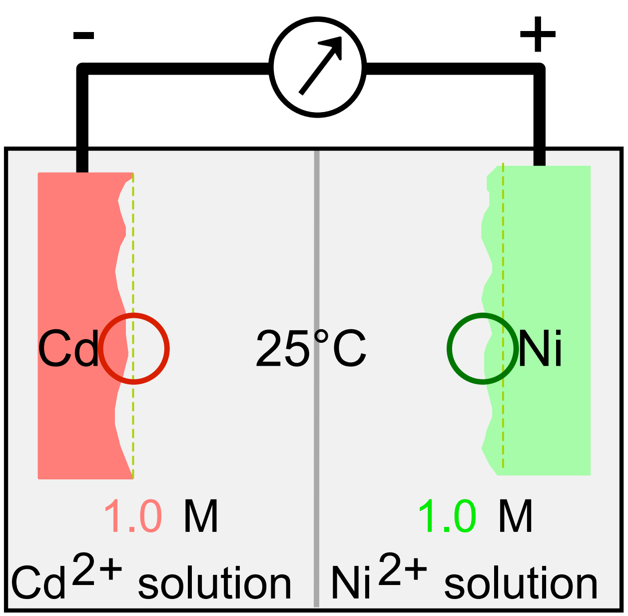


Adapted from Fig. 17.2, *Callister 7e.*

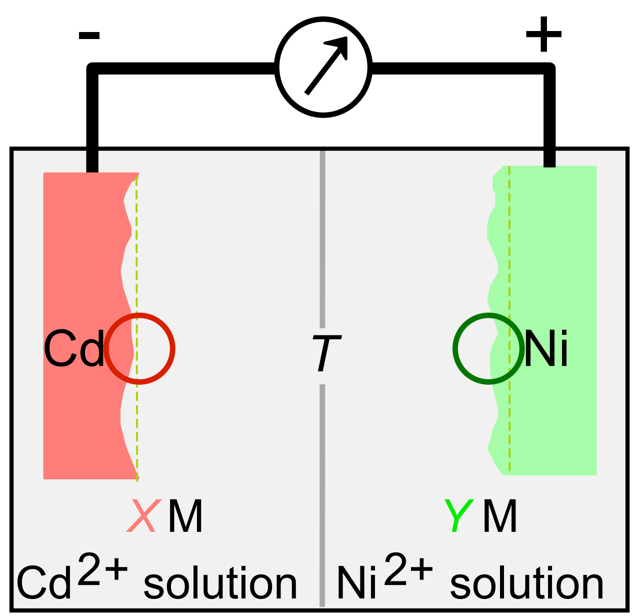
Effect of solution concentration

- Ex: Cd-Ni cell with standard 1 M solutions

$$V^{o}_{Ni} - V^{o}_{Cd} = 0.153$$



- Ex: Cd-Ni cell with non-standard solutions



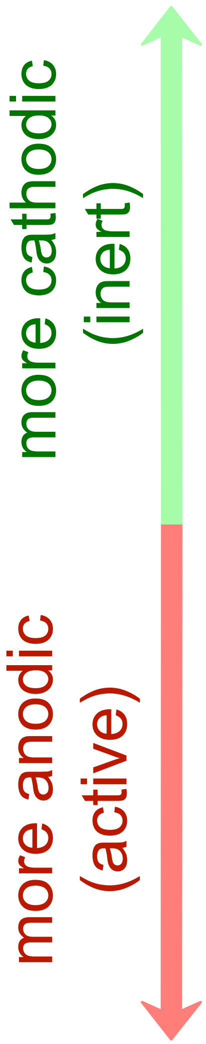
n = #e$^-$ per unit oxid/red reaction (= 2 here)

F = Faraday's constant = 96,500 C/mol.

- Reduce V_{Ni} - V_{Cd} by
  - --increasing X
  - --decreasing Y

Galvanic series

- Ranks the reactivity of metals/alloys in seawater



Platinum
Gold
Graphite
Titanium
Silver
316 Stainless Steel
Nickel (passive)
Copper
Nickel (active)
Tin
Lead
316 Stainless Steel
Iron/Steel
Aluminum Alloys
Cadmium
Zinc
Magnesium

Based on Table 17.2, *Callister 7e*. (Source of Table 17.2 is M.G. Fontana, *Corrosion Engineering*, 3rd ed., McGraw-Hill Book Company, 1986.)

Forms of corrosion

- **Stress corrosion**
Stress & corrosion work together at crack tips.

- **Uniform Attack**
Oxidation & reduction occur uniformly over surface.

- **Erosion-corrosion**
Break down of passivating layer by erosion (pipe elbows).

- **Selective Leaching**
Preferred corrosion of one element/constituent (e.g., Zn from brass (Cu-Zn)).

Forms of corrosion

- **Pitting**
Downward propagation of small pits & holes.

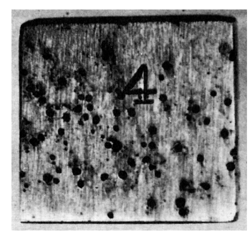


Fig. 17.17, *Callister 7e*. (Fig. 17.17 from M.G. Fontana, *Corrosion Engineering*, 3rd ed., McGraw-Hill Book Company, 1986.)

- **Intergranular**
Corrosion along grain boundaries, often where special phases exist.

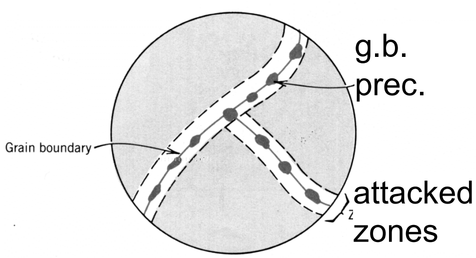


Fig. 17.18, *Callister 7e*.

- **Galvanic**
Dissimilar metals are physically joined. The more anodic one corrodes.(see Table 17.2) Zn & Mg very anodic.

- **Crevice** Between two pieces of the same metal.

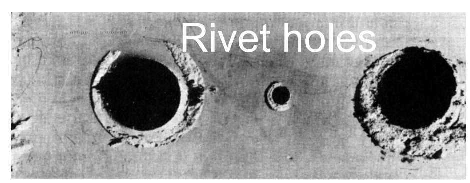
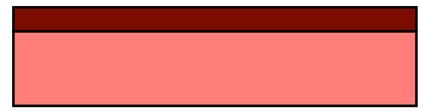


Fig. 17.15, *Callister 7e*. (Fig. 17.15 is courtesy LaQue Center for Corrosion Technology, Inc.)

Controlling corrosion

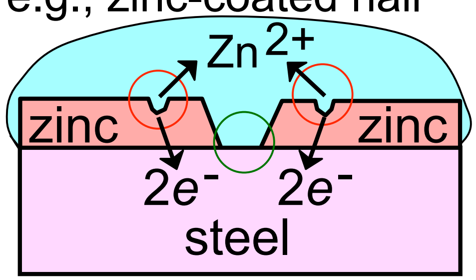
- Self-protecting metals!
  -- Metal ions combine with O
  to form a thin, adhering oxide layer that slows corrosion.



- Reduce *T* (slows kinetics of oxidation and reduction)
- Add inhibitors
  -- Slow oxidation/reduction reactions by removing reactants
  (e.g., remove O_2 gas by reacting it w/an inhibitor).
  -- Slow oxidation reaction by attaching species to
  the surface (e.g., paint it!).

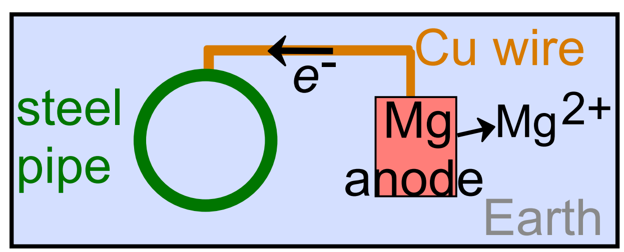
- Cathodic (or sacrificial) protection
  -- Attach a more anodic material to the one to be protected.

e.g., zinc-coated nail



Adapted from Fig. 17.23, *Callister 7e.*

e.g., Mg Anode



Adapted from Fig. 17.22(a), *Callister 7e*. (Fig. 17.22(a) is from M.G. Fontana, *Corrosion Engineering*, 3rd ed., McGraw-Hill Book Co., 1986.)